

# Synthesis and Application of Hematite from Lathe Waste as Anti-Swelling on Wood

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## ABSTRACT

Hematite is one of iron oxide, which potentially in reducing swelling wood. The cellulose content in wood easily binds with hydrogen that causing swelling. Hematite was synthesized from iron lathe waste via ferric precursor by a precipitation-calcination method. The product synthesized is characterized using X-Ray Diffraction (XRD), Scanning Electron Microscopy (SEM), Fourier Transform – Infrared Spectroscopy (FT-IR), color reader. The hematite pigmented wood powder analyzed swelling ratio to know the ability to reduce swelling. XRD data and FT-IR spectrum confirmed that after calcination obtained hematite. SEM image revealed the shape of hematite was spherical and uniform. Color reader data showed that product calcined got chroma, hue, lightness, degree of redness, and yellowness according to hematite standard. Swelling ratio analysis resulted that wood powder after interaction with hematite using H<sub>2</sub>SO<sub>4</sub> solvent did not swell, and using methanol solvent was swelled power after 8 hours of 4.76%.

**Keywords:** hematite, pigment, wood, anti-swelling

## 1. INTRODUCTION

Swelling in wood is a structural change caused by the relative humidity of the atmosphere [1]. Wood consists of cellulose around 38-51 %, glucomannan around 14-20 %, lignin by 21 – 31 %, xylan around 14-30 %, and other polysaccharide around 2-4 % [2], [3]. In the amorphous cellulose area, OH group arrangement is irregular to bind between hydrogen and water molecules [4]. According to Howsinon., J. A, the higher the degree of crystallinity of cellulose, the lower its water absorption. The degree crystallinity of wood pulp by 57% has to access to a formation with hydrate of 53%. In comparison, cotton with a degree of crystallinity of 67% has the power of formation with a hydrate of 44% [5].

One of the solutions to overcome swelling on wood is pigmented [6]. According to Schirp, Plinke, and Napolow, the brown iron oxide pigment from Bayferrox applied to wood-plastic composites can reduce its swelling power by about 44% compared to other dyes [7]. Crystallinity degree of cellulose increase with hematite pigment to suppress bonds with hydrates.

Iron oxide as a pigment because it has a variety of colors with different structures. Yellow is goethite; black is magnetite; brown is maghemite; red is hematite; and many others. Besides, iron oxide pigments are weather-

resistant, relatively inert, and nontoxic. Hematite iron oxide is also the most stable iron oxide phase among the others. Hematite is a red iron oxide pigment with the structural formula ( $\alpha$ -Fe<sub>2</sub>O<sub>3</sub>). Cornell and Schwertmann [8] explained that iron oxides could be synthesized from iron, Fe<sup>2+</sup>, or Fe<sup>3+</sup>. Hematite can be synthesized from pure compounds or iron lathe waste. Khoiroh, in his research, explained that the content of iron in lathe waste was 97 % [9].

Iron oxide pigments reduce the swelling of the wood. According to Lu., B [10], the hematite's surface is very reactive to water. The hematite surface can interact with water during dialysis, as evidenced by the FT-IR results, which show the number of OH groups. Composite hematite can reduce the surface area of the composite, thereby reducing its adsorption. Based on Ariyani [11], the composition 1:1 of iron-carbon oxide composite has a surface area of 219.59 m<sup>2</sup>/g while 0:1 a surface area of 1057 m<sup>2</sup>/g. Iron oxide composites: carbon has a smaller surface area than the surface area of carbon without being composited with iron oxide.

In view of hematite as a pigment, the present investigations were undertaken to synthesis hematite by the precipitation-calcination method. The precursors were prepared from lathe waste and followed the effect

of temperature. Hematite has been synthesized is applied to wood, and analyzed swelling on wood.

**2. EXPERIMENTAL**

**2.1 Material**

The materials used in this study were iron lathe waste, wood powder, methanol p.a, HNO<sub>3</sub> 65%, NaHCO<sub>3</sub> p.a, H<sub>2</sub>SO<sub>4</sub> p.a, were purchased from Merck.

**2.2 Method**

**2.2.1 making a precursor**

A total of 10 grams of iron lathe waste was put in a 500 mL glass beaker and added with 100 mL of 5 M HNO<sub>3</sub>. The solution allowed standing for 24 hours so that it dissolved completely then heated on a hot plate.

**2.2.2 Synthesis Hematite Pigment**

The precipitate is dissolved in 250 mL of distilled water. The mixture is stirred till it forms a slurry yellowish-brown mixture, then added 25 mL of 0.5 M NaHCO<sub>3</sub> and adjusted to pH 6. The mixture was then heated on a hotplate at 70°C and distilled for 1 hour, then cooled to room temperature. The mixture was decanted, then filtered with Whatman 41 paper and washed with 300 mL of distilled water. The precipitate is dried until it turns into powder. The powder is heated in a furnace at 750°C and 850°C for 3 hours and then cooled in a

desiccator to obtain a red powder [9]. The product was analyzed by XRD, SEM, color reader, and FT-IR.

**2.2.3 Pigmented wood with hematite iron oxide**

The resulting pigment is then dissolved with methanol at a ratio of 1: 10, stirring so that it is completely dissolved. Furthermore, the prepared wood powder interacted with the dissolved hematite pigment using a shaker for 20 minutes. The wood powder pigmented with hematite was filtered and oven at 30°C until constant weight, then analyzed by FT-IR. The same way is done using sulfuric acid, then tested its swelling ability at time intervals of 3, 8, 24, 48, and 72 hours.

**3. RESULT AND DISCUSSIONS**

**3.1 Characterization with XRF**

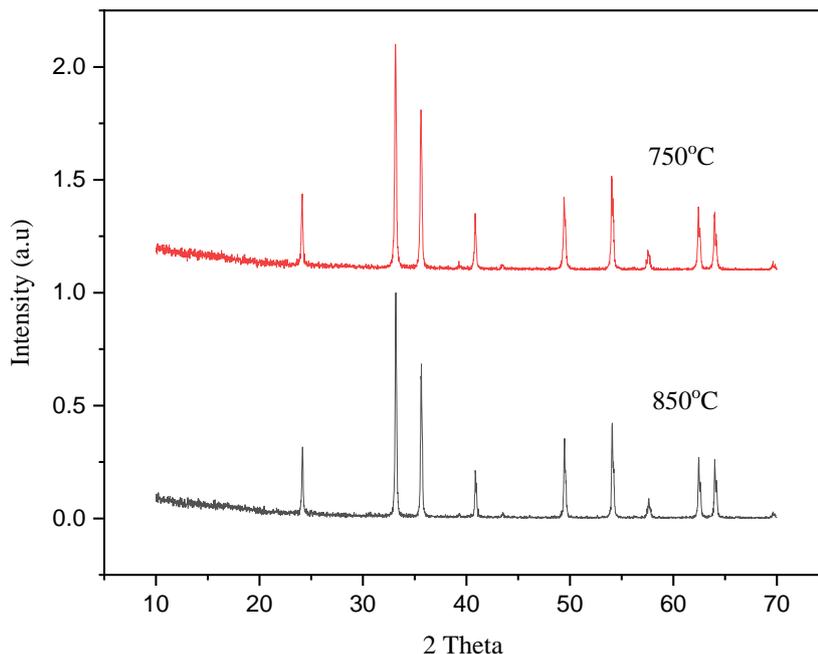
Based on the XRF results in Table 1, many elements found, but the number was small, except iron was 96.08 %. It was indicated that iron is the potential to be a precursor in synthesizing hematite.

**3.2 Characterization with X-ray Diffraction**

The X-ray diffraction (XRD) pattern of hematite is illustrated in Fig. 1. The figure shows that pure hematite achieved using the precipitation-calcination method at 750°C and 850°C.

**Table 1.** XRF data for raw material of lathe waste

Element	% Amount
Iron (Fe)	96.08
Mangan (Mn)	1.3
Europium (Eu)	0.9
Nickel (Ni)	0.82
Rubidium (Rb)	0.66
Calcium (Ca)	0.29
Chromium (Cr)	0.096
Cuprum (Cu)	0.093
Lanthanum (La)	0.02



**Figure 1.** Diffraction pattern of hematite

All of the peaks are indexed to rhombohedral  $\alpha\text{-Fe}_2\text{O}_3$  with space group R-3c and parameter  $a = 5.0356$   $c = 13.74$ . The peak can be attributed to the (012), (104), (110), (113), (024), (116), (018), (214), and (300) crystalline structures corresponding to pure  $\alpha\text{-Fe}_2\text{O}_3$  (matched with PDF No- 00-33-0664). No extra peaks for impurities were detected, indicating high purity and crystallinity in both diffractions. The higher crystallinity of hematite was achieved at 750°C, as shown by the higher intensity. The average crystallite size of hematite synthesized was calculated using Debye-Scherrer's equation. The average crystal size at 750°C was 0.15  $\mu\text{m}$ , and 850°C was 0.16  $\mu\text{m}$ . Both of temperatures were not significantly different, showing the temperature didn't affect the crystal size. Temperature changes that are not much different produce crystal size not different too.

**3.3 Characterization with Color Reader**

Based on the results of the color reader measurements in Table 2, the synthesized compounds after precipitation

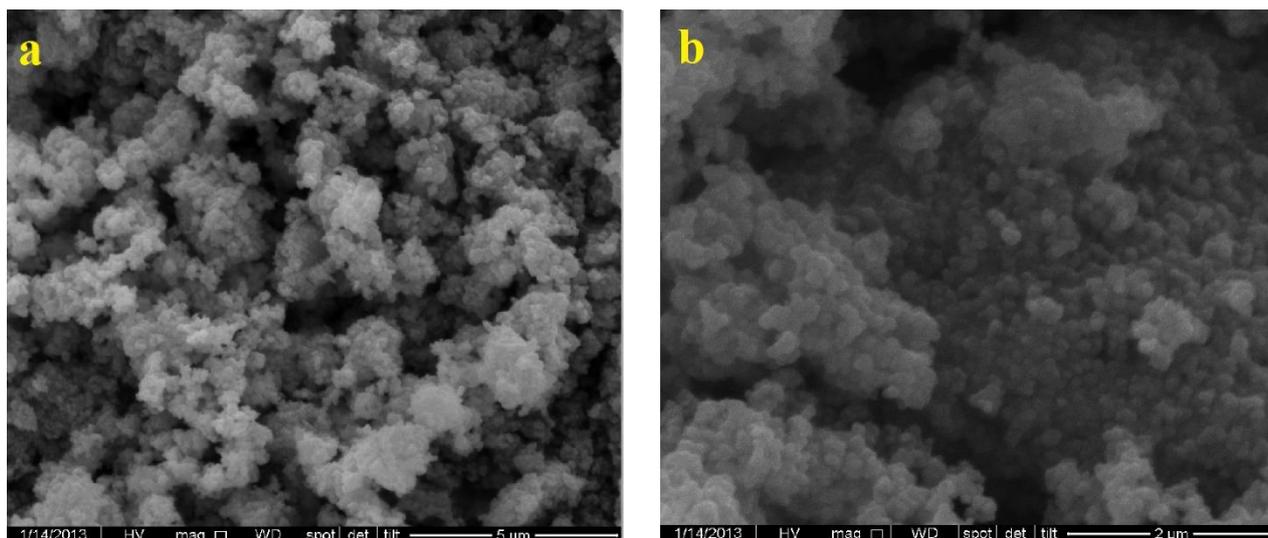
**Table 2.** Optical data of product synthesized

Condition	Temperature of synthesis	Color				
		L*	C*	H°	a*	b*
After Precipitation	70°C	37,2	23,5	17,6	29,36	36,8
After Calcination	750°C	33,5	20,4	35,0	16,7	12,3
	850°C	32,5	19,2	30,7	16,5	9,8

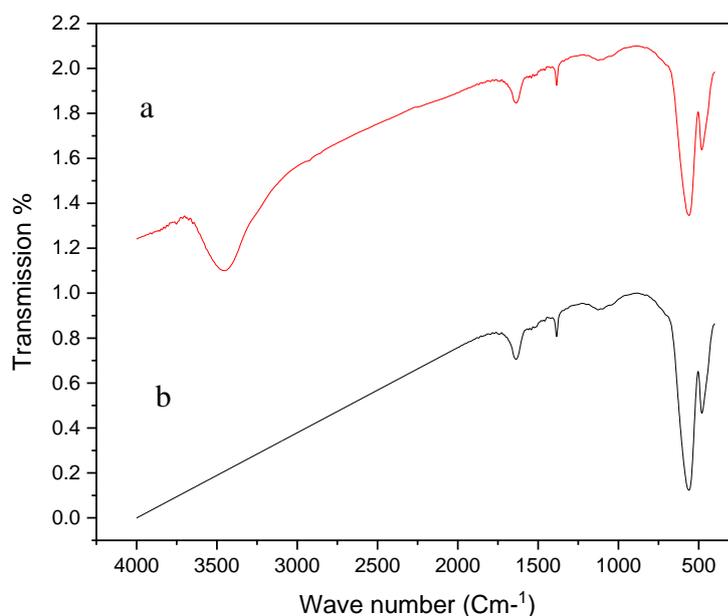
was by akagenite standard except for hue value and after calcination was following the standard. At 750°C, the degree of yellowness, redness, hue, and saturation was higher than 850°C; this is due to the larger grain size at that temperature.

**3.4 Characterization with SEM**

The morphology of hematite at 850°C revealed particle shape distribution is relatively spherical and uniform. The particle size of hematite at 750°C is around 0.09  $\mu\text{m}$ , as seen in Figure 2. The shape, size, and distribution of particles are the factors that influence the resulting color. Particle size and distribution of particles cause different color values in the optical analysis. The degree of redness, yellow, hue, and chroma hematite 750°C is greater than the synthesized hematite at a temperature of 85°C.



**Figure 2.** SEM image of hematite at 850°C: (a) low magnification and (b) higher magnification



**Figure 3.** FT-IR spectrum of product synthesized; (a) after precipitation (b) after calcination

### 3.5 Characterization with FT-IR

Figure 3 shows the FTIR spectra of iron oxide. All of the vibration after precipitation same with calcination except at  $3448\text{ cm}^{-1}$ , indicating OH vibration. There is OH-bending vibration at  $1636\text{ cm}^{-1}$ , there is  $\tau$ -O vibration in the bond plane at  $1384\text{ cm}^{-1}$ , and the t-O vibration at  $560\text{ cm}^{-1}$ , and the Fe-O stretching vibration at  $499\text{ cm}^{-1}$ . Based on these vibrations, the possibility of a product after precipitation is akagenite.

The characterization of swelling power was carried out to determine synthesized compounds' ability to reduce wood swelling. Wood is a material that contains high levels of cellulose, lignin, hemicellulose, and other contents. One solution to avoid the problem of fungus on wood is to coat it with pigments.

The wood interacted with the synthesized compound, dissolved in two different solvents, namely using  $\text{H}_2\text{SO}_4$  and methanol. After interacting with the dissolved pigment, then dried and characterized by FT-IR, as shown in Figure 4. The spectra show that hematite pigmented wood with methanol solvent revealed OH

vibrations at around  $3427\text{ cm}^{-1}$ , C=O vibration at about  $1623\text{ cm}^{-1}$ , C=C at around  $1509\text{ cm}^{-1}$ , and vibrations of Fe-O stretching at  $485\text{ cm}^{-1}$ . Spectra IR hematite pigmented wood with  $\text{H}_2\text{SO}_4$  solvent is not much different from methanol solvent; just vibrations for  $\text{H}_2\text{SO}_4$  solvent has higher intensity.

### 3.6 Swelling Ratio of Wood Powder

The wood powder interacted with the pigment is analyzed by swelling using water. Based on Figure 5, the wood powder interacting with hematite using  $\text{H}_2\text{SO}_4$

solvent did not swell at all. If with methanol solvent, the swelling power of wood powder is 4.76% at 8 hours. The wood not interacted with the hematite has experienced swelling of 2.5% at 3 hours, and the swelling power increases with time.

The wood's swelling shows that the synthesized compound can coat wood powder to prevent swell on wood. Pigments can prevent hydrogen bonding with cellulose in wood, thereby minimizing hydrogen bonds, causing swelling.

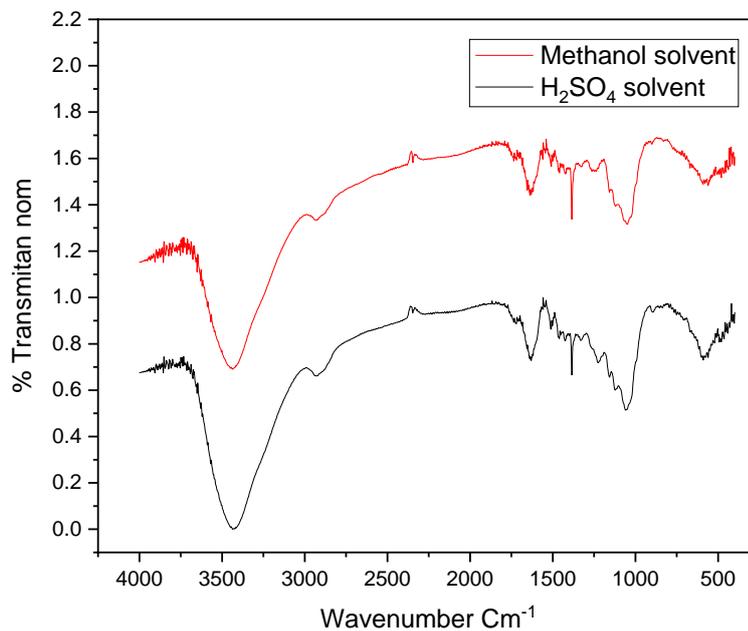


Figure 4. FT-IR spectra of hematite pigmented wood

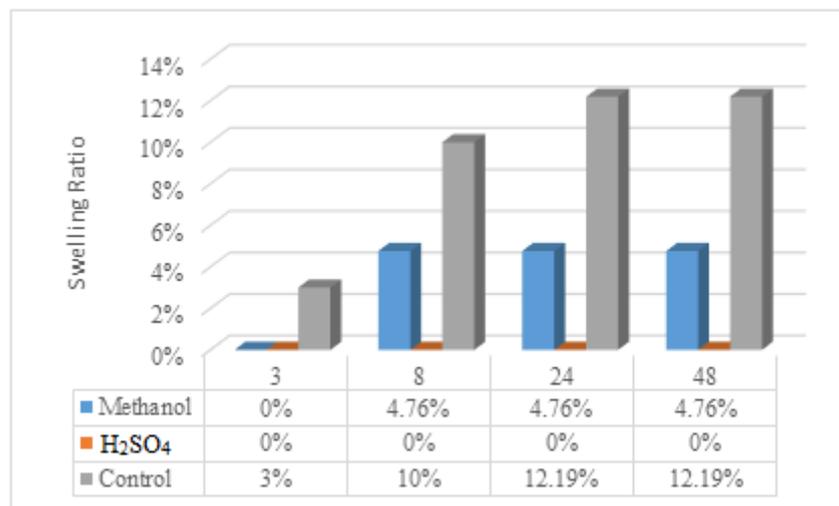


Figure 5. Graph of the swelling ratio of hematite pigmented wood powder

#### 4. CONCLUSION

Hematite with a high degree of crystallinity can be synthesized from lathe waste at 750°C. Hematite pigment can coat the wood to prevent swell. Hematite pigmented wood using H<sub>2</sub>SO<sub>4</sub> solvent didn't experience swelling at all.

#### ACKNOWLEDGMENTS

Some of this research are funded by Science and Technology Faculty UIN Maulana Malik Ibrahim Malang.

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