

The Controllable Synthesis of Silver Nanowires by Using the Polyol Method with Trace Agent

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Abstract. Ag nanowires were successfully synthesized by the polyol process with the addition of trace amounts of HCl and HNO₃. In addition, the relationship between the final morphology of the silver products and the reaction parameters, such as the quality ratio of HCl and HNO₃ were studied detailed. The results show that when the quality ratio of HCl to HNO₃ is 1:2, almost all of the silver products were thin nanowires with uniform size 40 to 60 nm in diameter and 30 μm in length.

1 Introduction

Transparent electrodes are critical for many electronic devices such as liquid crystal displays, touch screens, plasma displays, organic light emitting diodes, and solar cells [1,2]. Transparent electrodes with random networks of silver nanowires (AgNWs) exhibit high electrical conductivity and transparency, because silver has the lowest resistivity among metals. Ag nanowires (NWs) are mostly synthesized by a polyol reduction method with poly(vinylpyrrolidone) (PVP) as the capping agent[3]. Xia et al. reported the possibility of controlling the morphology of nanomaterials under the existence of PVP(capping agent)by controlling the reaction conditions, such as the molar ratio of the reactants, reaction temperature, reaction time, in ethylene glycol (EG)[4].

Mao and co-workers carefully studied Xia's self-seeding method and demonstrated that the selective adsorption of PVP on the side surfaces(100) and end surfaces(111) plays a crucial role in the growth of Ag NWs[5]. In addition, various additives, such as FeCl₃[6], CuCl₂[7], Na₂S[8], NaCl[9], PtCl₂[10], KNO₃[11], AgBr [12], have been commonly employed and determined to play an essential role in the selective formation of Ag NWs. The cations or anions from these additives are crucial for the seed formation and uniaxial growth of multiple twinned particles for the polyol synthesis of Ag NWs[6,7]. Without fine control of reactant concentrations, especially the additives, some by-products, such as nanocubes or nanospheres, can be produced from the solution. Recently, we have sought to expand on this work by looking at the effect of Cl⁻ and H⁺ on the polyol reduction of AgNO₃. In this work, we synthesized Ag NWs with polyol method with trace amounts of HCl and HNO₃ to controllably adjust the diameter and length of the Ag NWs. We found that the addition of both HCl and HNO₃ facilitated the growth of Ag nanowires.

2 Experimental Section

2.1 Material

Poly(Nvinylpyrrolidone) (PVPK 30, Mw~40,000, the concentration was calculated in terms of the repeating unit) was purchased from Shnopharm Chemical Reagent

Co. Ltd. Silver nitrate (AgNO₃, 99.9%), ethylene glycol (EG), diluted hydrochloric acid (HCl 1.8%w) and diluted nitric acid (HNO₃ 3.2%w) were purchased from Sinopharm Chemical Reagent Co. Ltd. All chemicals were of analytical grade and were used without further purification.

2.2 Method

Silver nanowires were synthesized by reducing AgNO₃ with EG in the presence of PVP. A 50 mL EG solution of 0.09M AgNO₃ was preheating at 120°C for 30min. Another 50 ml EG solution of 0.54M PVP with trace amounts of HCl and HNO₃ was injected drop by drop using a syringe into the preheated solution. The mixture solution was maintained at 142°C for 1h.

For these syntheses that produced a high-yield of Ag nanowires under protection of argon, the following color change were observed: initially clear and colorless to light yellow (within 2 min), to orange-red (5 min later), beginning to green (with 20 min), and finally to opaque gray with wispieness indicating the formation of long nanowires (within about 1h). Upon nanowires formation, the reaction was naturally cooled to room temperature (25°C). Silver nanowires were easily separated from EG by addition of a large amount of acetone and ethanol, followed by sonication and centrifugation. Then the Ag NWS products were dispersed in ethanol for further characterization.

2.3 Characterization

The morphology of the synthetic Ag NWs was analyzed by field emission scanning electron microscopy equipped with energy dispersive X-ray spectroscopy (EDX) (FESEM, S-4800, Hitachi, Japan) operated at an acceleration voltage of 5 kV. Powder X-ray diffraction (PXRD) patterns were recorded on an X-ray diffractometer (Shimadzu XRD-7000) using Cu Kα radiation with a scan rate of 2θ = 4°/min and a range of 2θ = 25~85°.

3 Results and Discussion

The silver nanomaterials were obtained by reducing AgNO_3 with EG in the presence of PVP. Besides, trace additives of HCl and HNO_3 were added into the reaction system to investigate its influence on the morphology of the silver nanoproducts. In order to examine the presence of Ag nanowires, the typical XRD pattern of as-prepared products is shown in Fig.1. Four characteristic diffraction peaks are observed. The characteristic diffraction peaks positioned at 37.9° , 44.1° , 64.2° , 77.3° , and 81.53° are ascribed to (111), (200), (220), (311) and (222) planes of face-centered cubic (fcc) silver (JCPDS, No.04-0783), respectively, which shows that the products were Ag.

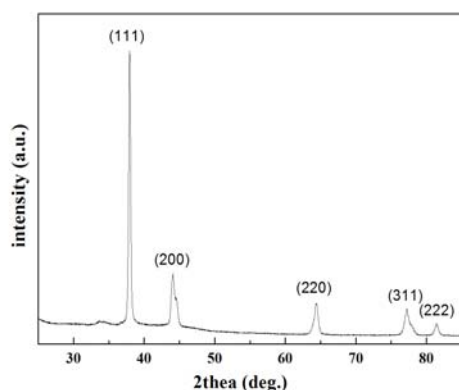


Fig 1. XRD pattern of AgNWs.

PVP is well-known to selectively passivate on {100} planes of Ag seeds, leaving the unpassivated {111} planes to grow longitudinally, but a little HCl and HNO_3 also play a necessary role on uniform growth of silver. Fig 2 shows typical scanning electron microscopy (SEM) images of the as-obtained Ag nanocrystals based on different growth conditions. The most of samples without addition of HCl and HNO_3 present a morphology of nanowire, but there is still a certain proportion of particles, while Ag NWs had shorter length ($2\sim 5\mu\text{m}$) and larger diameter with about $95\sim 120\text{nm}$ (as shown in Fig 2a). With the addition of hydrochloric acid, the nanowire samples are distinctly observed (in Fig 2b), only little particles existed. However, when the quality ratio of HCl to HNO_3 is 1:2, almost all of the silver products are thin nanowires (Fig 2c) with nearly uniform size, measured 35 to 50 nm in diameter and length up to $35\mu\text{m}$. Under the condition of only the presence of HNO_3 , a small amount of nanowires gradually disappear, instead of the larger particle products, and the length to diameter ratio of nanowires declined dramatically (Fig 2d). We attribute this difference to the absence of Cl^- , which is needed to control the solution saturation when reagents are added rapidly by pipette.

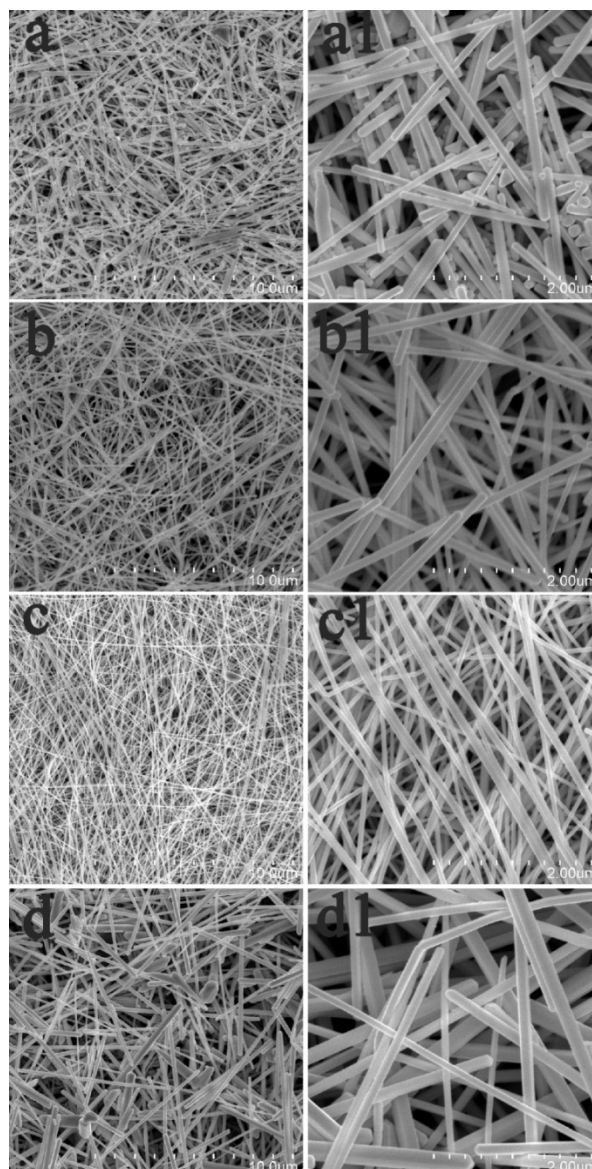


Fig 2. FESEM images with high magnification (a-d) and low magnification (a1-d1) of Ag NWs synthesized at different quality ratios of HCl to HNO_3 : a. (0g HCl , 0g HNO_3); b. (0.05g HCl , 0g HNO_3); c. (0.05g HCl , 0.1g HCO_3); d. (0g HCl , 0.1g HNO_3)

Fig 3 shows the scattergrams of the diameter and length of Ag NWs under the different trace additive agent of HCl and HNO_3 . We found that the addition of both HCl and HNO_3 facilitated the growth of Ag nanowires, which not only affects the length and the diameter of the nanowires, but the uniformity of the AgNWs. When the mass ratio of HCl to HNO_3 is 1:2, the ratio of length to diameter of Ag NWs is the largest. Consequently, at the same time of Ag nanocrystals grow preferentially along these facets due to PVP selected growth, trace control agent and reaction conditions, such as pH, can also impact the morphology of silver nanoproducts.

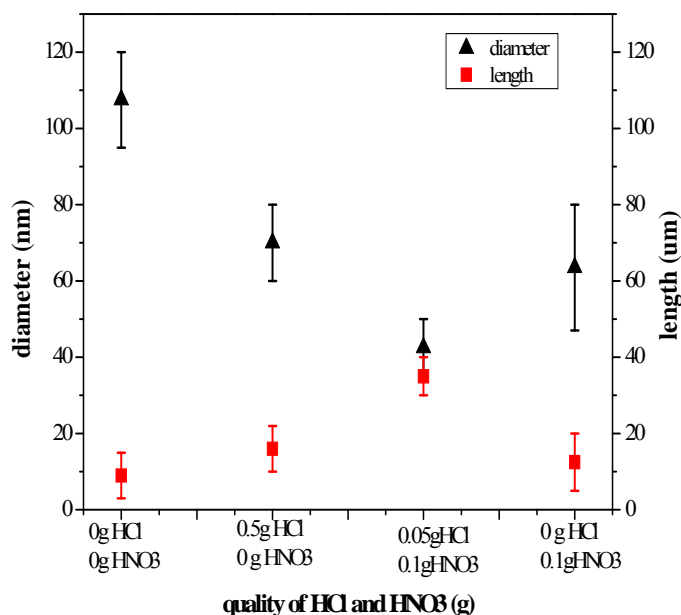


Fig 3. Distribution of diameters and lengths for the batch of AgNWs shown in the SEM images of Fig 2.

At the initiation of the reaction without adding hydrochloric acid, because of the more free silver ion in the solution, the faster reaction rate lead to product as isotropic silver particles. Once hydrochloric acid added, AgCl precipitation formed immediately, which then reduced the nucleation rate of silver. As the reaction progresses, the reaction rate decreased, products began to anisotropic growth. This is mainly because: 1. electrical activity of the oxidation of AgCl/Ag is much lower than that of Ag⁺/Ag, the formation of chelate AgCl increased the stability of silver ions; 2. ethylene glycol is very weak reductant, and joined drop by drop, which in not enough into the reaction system; 3. The concentration of silver ion in solution reduce a lot during the reaction process.

Fig 4 exhibits the growth process of silver nanowire. growth process of wires includes three aspects: (a). isotropic nucleation formation of silver nanoparticles; (b). anisotropic growth process on the surface of nanoparticles makes Ag crystals develop in length to diameter; (c). elemental silver reduced concentrate on the lower surface, eventually forming the nanowires.

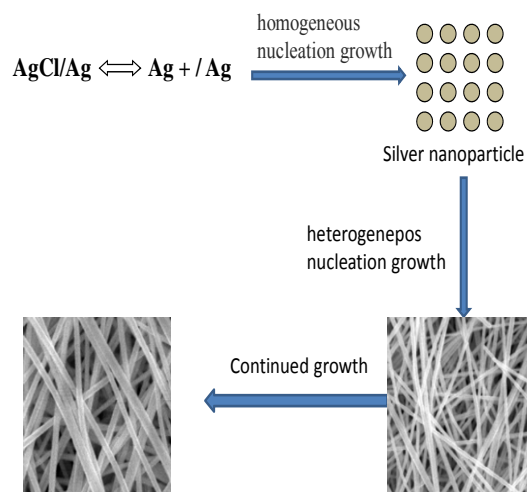


Fig 4. The growth process of Ag NWs

4 Conclusions

Ag NWs were successfully synthesized by this self-seeding process through a polyol method. It is found that the addition of trace amounts of HCl and HNO₃ greatly facilitates the formation of silver nanowires. Besides, the relationship between the final morphology of the products and the quality ratio of HCl to HNO₃ was studied. Cl⁻ combines Ag⁺ to form amounts of AgCl colloids in the initial stage, resulting in decreasing free Ag⁺ during initial formation of silver seeds and slowly releasing of Ag⁺ to the solution in the subsequent reaction. The formation of chelate AgCl increased the stability of silver ions. The concentration of H⁺ play an essential role in the selective formation of Ag nano-products, mainly affected greatly the electrical activity of AgCl/Ag to Ag⁺/Ag, which then is effective in reducing the nucleation rate of silver. Finally, nanocrystal began to anisotropic grow to be AgNWs. In addition, silver nanowires with controllable sizes can be obtained by adjusting the concentration of HCl and HNO₃.

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