

Synergetic Effect on Syngas Yield and Energy Yield during Co-pyrolysis of Sewage Sludge and Rice Straw

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Abstract. This study focused on the co-pyrolysis of sewage sludge and rice straw in order to improve syngas yield and energy recovery potential. The singular pyrolysis of sewage sludge and rice straw was firstly investigated under varying operating temperatures. Then the synergetic effect on yield of combustible gas and energy yield during co-pyrolysis was studied. Results showed that the syngas yield increased with operating temperature from pyrolysis of sewage sludge, meanwhile, the energy recovery reached maximum value of 5194.3 J·g⁻¹ at 850°C. The yields of combustible gases were found to first increase and then decrease with increasing pyrolysis temperature from pyrolysis of rice straw. The synergetic effect of co-pyrolysis on gas yield was different at different operating temperature. There was a beneficial synergetic effect on yields of CO and H₂ at 850°C, while it was shown an unfavorable influence on CO and H₂ yields during co-pyrolysis at 650°C. The maximum conversion of energy reached 8098.4 J·g⁻¹ at 850°C when the weight ratio of rice straw was 15%. In this case, it was indicated that an energy recovery of approximately 60% was achieved with the energy potential of parent stock.

1 Introduction

Municipal wastewater treatment results in the generation of excessive amount of sewage sludge (SS), so that the disposal of SS is a matter of growing concern. Conventional SS disposal methods, including using as soil conditioner in agriculture, landfills and ocean dumping, have shown some disadvantages because of negative environmental impacts and energy waste [1,2]. In recent years, a more environmentally friendly and renewable energy alternative such as thermochemical conversion of SS has been turned out. In the pyrolysis process, the SS is thermally decomposed into the products of syngas, bio-oil and char in an anoxic atmosphere. Energy recovery in form in syngas and bio-oil could be achieved with the energy potential of the original SS[3].

Many researchers have studied the effects of pyrolysis parameters on the yield and composition of the derived products from dried SS. Especially, more and more research has been focused on the conversion of SS into bio-oil production as a high-energy feedstock[4,5]. But to the syngas production, there are limited research on gas yield, composition and heating value because of its relatively lower heating value[6]. However, syngas production with advantages of clean fuels, easier to collect and transport plays an important role in energy utilization. In order to improve the combustible gas yield and energy recovery potential, some researchers focused on co-pyrolysis of sewage sludge and biomass[7,8]. It was indicated that synergetic effects occurred during co-pyrolysis process. The alkali and alkaline earth metals provided by SS and biomass were considered to act as catalyst and facilitate H₂O-char and CO₂-char gasification which reduce the H₂O and CO₂ yields and benefit the increase of H₂ and CO yields[7]. However, disputed results were obtained in different pyrolysis conditions including biomass species and ratio, pressure, operating temperature and operation mode[9-11].

In this study, rice straw (RS) which is a valuable biomass feedstock was selected as additive because of its abundance. The singular pyrolysis of sewage sludge and rice straw was firstly investigated under varying operating temperatures in a fixed bed reactor. Then the synergetic effect on syngas yield and energy yield during co-pyrolysis of SS and RS was studied. This study focused on the maximum of energy recovery in form in syngas from the co-pyrolysis process, by optimizing operating temperature and ratio of rice straw.

2 Material and methods

2.1 Materials

The sewage sludge (SS) was supplied from Xianyang Road Municipal Wastewater Treatment Plant in Tianjin. The moisture content of the as-received SS was approximately 85 wt.%. Prior to use, the SS was air-dried for several days to remove the unbound moisture. The rice straw (RS) was originated from rural area of Tianjin. The air-dried SS and RS were sequentially dried in an oven for 12 h at 105±2°C. Then they were ground into small particles and screened with 0.25-0.38 mm.

The proximate and ultimate analyses were taken place in a thermogravimetric analyzer and element analyzer respectively according to GB483-87 in China as table 1 shown. Contrasting with sewage sludge, the rice straw contained a high content of volatile which was up to 80%. The ash content of the SS was found to be 39.1 wt% which was much higher than that of RS. The lower heating value (LHV) of dried SS and RS was 11.2 and 27.1 MJ·kg⁻¹ respectively.

Table 1. Analysis of raw material.

Proximate analysis (wt%)			Ultimate analysis (wt%)		
	SS	RS		SS	RS
Moisture	6.5	1.0	C	22.0	39.5
Volatile	48.5	80.1	H	2.1	5.7
Fixed carbon	5.9	11.1	O	23.9	29.3
Ash	39.1	7.9	N	3.5	0.8
LHV(MJ·kg ⁻¹)	11.2	27.1	S	2.9	1.1

The blends as parent stock containing varying 5 wt.% to 25 wt.% RS were used for studying the synergetic effect on gas yield and energy yield during co-pyrolysis process.

2.2 Experimental apparatus and procedure

The pyrolysis experiments were carried out in a fixed bed system with approximately 10 g samples used in each experiment. The pyrolysis system consisted of three parts: pyrolysis reactor, heater and condenser. The length of the quartz tube reactor with constant temperature was 400 mm and the inside diameter was 50 mm. The each end of reactor were fixed and connected with screw joint which welded to a 6 mm pipe. The inlet was connected to a nitrogen cylinder and the outlet was connected to the ice bath condenser, respectively.

Pyrolysis experiments were carried out in a nitrogen atmosphere prevailing in the pyrolysis reactor. The nitrogen was pumped into quartz tube reactor with 300 ml·min⁻¹ during the pyrolysis process. The reactor was heated at a rate of 10 °C·min⁻¹, and when the required temperature was attained, the stock in a corundum crucible (20 cm × 4 cm) was put into quartz tube reactor quickly. It was maintained until no pyrolysis gas was generated. The pyrolysis of stock was evaluated at various temperatures (450, 550, 650, 750 and 850 °C). The gases produced from the pyrolysis process were firstly passed through the ice bath condenser to eliminate the tar, the remaining gases were pumped into the gas analysis meter. The percentage contents of H₂, CO, CO₂, CH₄ and C_nH_m can be real-time measured. The yields of gas produced were calculated by integration method.

2.3 Calculations

To clearly determine the gas yield change between experimental value obtained from co-pyrolysis and theoretical value, the relative deviation (RD) was calculated by Eq. 1. The average relative deviation was calculated by Eq. 2.

$$RD = \frac{|\text{Experimental value} - \text{Theoretical value}|}{\text{Theoretical value}} \times 100\% \quad (1)$$

$$ARD = \frac{\sum_i^n RD_i}{n} \quad (2)$$

The energy yields (EY) in form in syngas during pyrolysis of SS, RS and blends were calculated by Eq. 3.

$$EY = \frac{\text{LHV of gas products}}{\text{LHV of parent stock}} \times 100\% \quad (3)$$

3 Results and discussion

3.1 Yield of gaseous products from SS and RS singular pyrolysis

Trends in the gas yield from SS and RS singular pyrolysis at different temperatures were illustrated in Fig.1. From the results presented in Fig. 1(a), it can be seen that temperature had a marked influence on the gas yields from pyrolysis of SS. The H₂ and CO yields were found to improve proportionally with pyrolysis temperature. Maximum yield of H₂ and CO reached 109.8 ml·g⁻¹ and 45.1 ml·g⁻¹ at 850°C respectively. The increased CH₄ yield obtained with increasing pyrolysis temperature from 450 to 650°C. Moreover, it remained relatively constant at 650, 750 and 850°C. The CO₂ yield decreased with increasing pyrolysis temperature. The significantly reduced yield of CO₂ obtained at 750 and 850°C. It was highly possible that CO₂-char gasification was intense in the high pyrolysis temperature, which reduced the CO₂ yield and benefited the increase of CO yield.

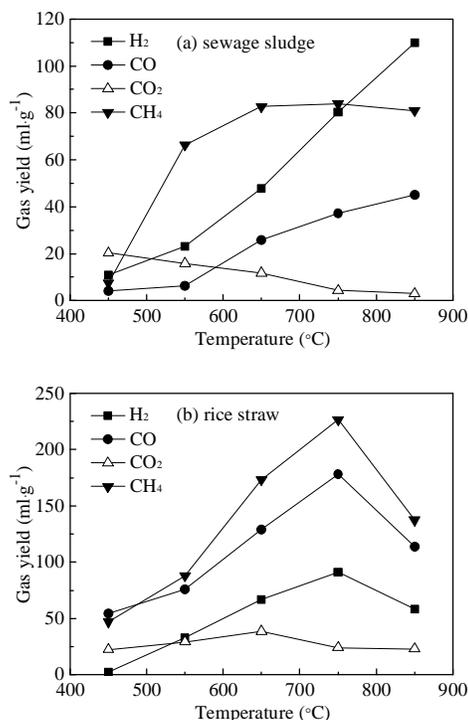


Figure 1. Gas yields of SS and RS singular pyrolysis at different temperatures.

The temperature also had a significant influence on the gas yields from pyrolysis of RS as Fig. 1(b) shown. The yields of combustible gases, including H₂, CO and CH₄, were found to first increase and then decrease with increase of pyrolysis temperature. The maximum yield of CH₄, reached 226.6 ml·g⁻¹ at 750°C, was evidently higher than the maximum value from SS pyrolysis. The CO production was also higher than that produced from SS pyrolysis, and the maximum yield of CO was obtained at 750°C. In contrast, the quantity of H₂ originated from RS pyrolysis was much lower than that of SS pyrolysis at moderate and high temperature.

3.2 Yield of gaseous products from SS and RS co-pyrolysis

When the co-pyrolysis process carried out at operating temperature of 850°C, the experimental and theoretical gas yields of pyrolysis blends containing varying weight ratios of RS were illustrated in Fig. 2. The experimental CH₄ yields of pyrolysis blends were obviously higher than the calculated values as Fig. 2(a) shown. Besides, the experimental yield of CH₄ first increased and then decreased with the increasing ratios of RS in blends. The maximum yield of CH₄ reached 140.9 ml·g⁻¹ when the weight ratio of RS was 15%. In this case, the relative deviation between the experimental CH₄ yield and theoretical value was 59.3%. The experimental yield of H₂ from pyrolysis of blends was also higher than the calculated values, except for blends containing 25 wt.% RS. Trends in H₂ yield from blends pyrolysis, either experimental value or theoretical value, was decreased with increasing ratio of RS.

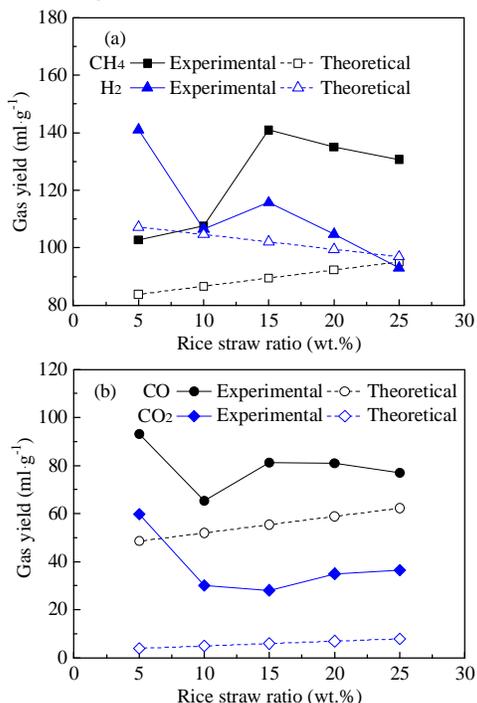


Figure 2. Theoretical and experimental gas yield of pyrolysis blends containing varying weight ratios of rice straw (at 850°C).

The CO yield from blend pyrolysis were higher than calculated value varying weight ratios of RS, as Fig. 2(b) shown, and the average relative deviation was 45.1%. The CO₂ yields from pyrolysis blends were also higher than the calculated values. Moreover, the changing trend of experimental value was almost parallel with trend of calculated value varying from 10 wt.% to 25 wt.% RS. The results above indicated that there were beneficial synergetic effects on yield of syngas during co-pyrolysis of SS and RS at operating temperature of 850°C.

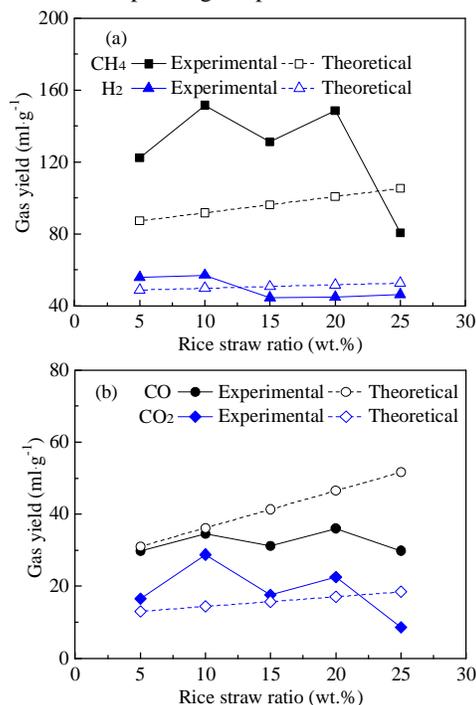


Figure 3. Theoretical and experimental gas yield of pyrolysis blends containing varying weight ratios of rice straw (at 650°C).

When the operating temperature remained at 650°C, the theoretical and experimental gas yield of pyrolysis blends containing varying weight ratios of RS were illustrated in Fig. 3. The CH₄ yields from co-pyrolysis were obviously higher than the calculated values as Fig. 3(a) shown. The average relative deviation of CH₄ yields reached 42.5%. The experimental H₂ yields were relatively close to the calculated values. H₂ production originated from blends pyrolysis was higher than calculated value when the weight ratios of RS was lower than 15%, while H₂ yield was lower than the calculated value when weight ratios of RS increased.

The contrary result was obtained for CO yield at operating temperature of 650°C contrasted with the result obtained at 850°C. The CO yield from co-pyrolysis were found to be lower than theoretical value varying weight ratios of RS at 650°C as Fig. 3(b) shown. The average relative deviation between CO yield and theoretical value was 19.6%. The experimental yield of CO₂ was slightly higher than the calculated values, except for blends containing 25 wt.% RS.

According to results in Fig. 2 and Fig. 3, it can be found that the synergetic effect occurred during co-pyrolysis process. However, the synergetic process had a

different influence on gas yield at different operating temperature. There was a beneficial synergetic effect on yields of CO and H₂ during co-pyrolysis of SS and RS at 850°C, while it was shown an unfavorable influence on CO and H₂ yields during co-pyrolysis at 650°C.

3.3 Energy recovery of syngas from SS and RS co-pyrolysis

Fig. 4 illustrated the experimental and theoretical lower heating value (LHV) of syngas from co-pyrolysis of SS and RS.

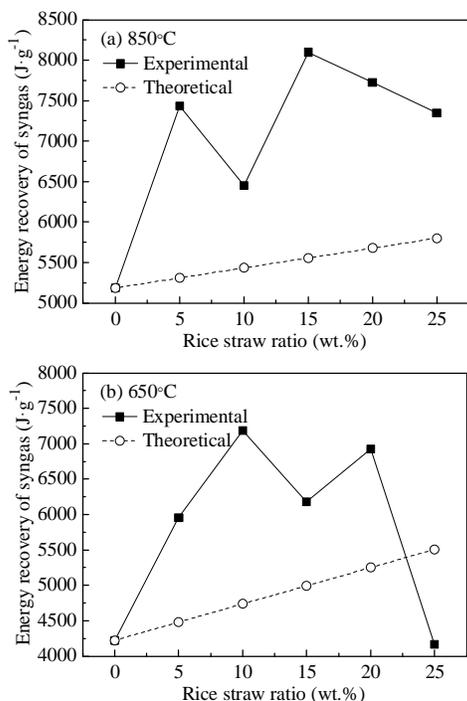


Figure 4. Theoretical and experimental energy recovery of syngas during co-pyrolysis.

As Fig. 4(a) shown, at operating temperature of 850°C, the conversion of energy from blends to the syngas production during co-pyrolysis were evidently higher than calculated values varying weight ratios of RS because of significantly increase of combustible gas yields, including CH₄, H₂ and CO yield. The average relative deviation between energy recovery of syngas and calculated value was 33.5%. The maximum conversion of energy from parent stock to syngas reached 8098.4 J·g⁻¹ when the weight ratio of RS was 15%. In this case, it was indicated that an energy recovery of approximately 60% was achieved with the energy potential of the original SS and RS of 13585 J·g⁻¹.

The conversion of energy from blends to the syngas production at 650°C were evidently lower than that obtained at 850°C. As fig. 4(b) shown, at operating temperature of 650°C the energy of syngas produced during co-pyrolysis were higher than calculated values because of higher CH₄ yield except for blends containing 25 wt.% RS. The maximum conversion of energy reached 7184.5 J·g⁻¹ in condition of blends containing 10 wt.% RS,

and LHV of syngas reached 56.2% of energy potential of the parent stock.

4 Conclusions

The obviously synergetic effects on syngas yield and energy yield were observed during co-pyrolysis of SS and RS. There was a beneficial synergetic effect on yields of CO, H₂ and CH₄ from blends pyrolysis at 850°C. The maximum conversion of energy from parent stock to syngas was achieved approximately 60% of energy potential of parent stock. It was shown an unfavorable synergetic effect on CO and H₂ yields during co-pyrolysis at 650°C. Whereas, heating values of syngas were higher than calculated values except for blends containing 25 wt.% RS because of higher yield of CH₄.

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