

## **Preparation and characterization of nano-structured tungsten carbide crystallites\***

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With  $WCl_6$ ,  $C_2H_2$  and  $H_2$  reaction system and Ar work gas, WC nanocrystallites were prepared by plasma chemical vapor deposition (PCVD). The properties of  $WCl_6$  and WC particles including chemical composition, morphology, size and structure of crystal phase were characterized by Transmission Electron Microscopy (TEM) and X-ray Diffraction (XRD). Results show that the most particles of WC in the obtained nanocrystallites are spherical and the mean particle size is about 50nm.

*Keywords:* chemical vapor deposition; nano tungsten carbide; characterization.

### **1. Introduction**

New materials have been playing one of the most vital roles in modern science and technology development. As the forerunner of new materials, nano technology has been developing since 1980s as the focus of the research worldwide[1], and there into hard alloy with nano structure has been concentrated in the research for powder metallurgy[2-4]. Reports indicated that if the size of tungsten carbide(WC) crystalline particle was under  $1\mu m$ , the hardness and tensile strength would increase synchronously, and the even remarkable increase of them could be observed for the nanostructured particle[5]. WC as a hard alloy with high hardness and tensity displays excellent performance in cutting processing, thus breaking the traditional conceptions that hardness increase brought up tensile strength decline. Preparing nanoscale WC particle by PCVD, only simple operations are needed to obtain samples of good

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<sup>\*</sup>This work was supported by Natural Science Foundation of Shanghai (No. 16ZR1412400), the Innovation Key Program of Shanghai Municipal Education Commission (No. 13ZZ140), Natural Science Foundation of Shanghai (No. 15ZR1417100), Key Discipline Construction of Shanghai Polytechnic University (Materials Science and Engineering, XXKZD1601).

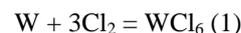
purity, thus creating an effective way to resolve the contradiction between hardness and tensile strength.

Solid phase reaction, liquid phase reaction and vapor phase reaction are the usual methods to prepare nano WC [6-9]. PCVD method is categorized as one way of the vapor phase reaction, which is characterized by high energy, fast reaction process, successive process and easy treatment with after-processing. Because reactants stay very short time in the plasma fire, they quickly cool down after leaving the fire, contributing to the small particle size, crystallization and sphere[10]. PCVD method is distinct for that it employs nonpolar discharge to prevent product from being polluted by melted or vaporized electrode at high temperatures. Also, it takes clean heat source and is easy to control. In this paper, nano WC microcrystal is prepared utilizing high frequency Ar as the heat source,  $WCl_6$  the tungsten source,  $C_2H_2$  and  $H_2$  the reaction gas source, and Ar the burning and cooling gas, obtaining WC nanocrystallite of average diameter about 50nm.

## **2. Experimental**

### **2.1. Preparation of $WCl_6$**

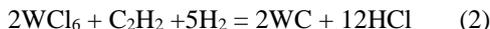
$WCl_6$  as the raw and processed material is difficult to buy for its scarcity. In our experiments,  $WCl_6$  was obtained by the reaction of  $Cl_2$  and W powder at 600~800°C via the reaction formula



Put the quartz tube containing W powder into electric stove, whose temperature was controlled by temperature controlling device connected to thermocouples, provide  $Cl_2$  and keep it in reaction for 3-4 hours. Asbestos were put into the joint of quartz tube and electric stove in order to keep heat from dispersing.

### **2.2. Preparation of WC**

WC was formed by vapor phase reactions in system  $WCl_6-C_2H_2-H_2$ , and the reaction was carried out in a high frequency plasma reactor using Ar as the fuel and cooling gas. Ar plasma of certain ionization constant was produced in the quartz tube thus creating high temperatures in the reactor of special structure. The mixed  $WCl_6$ ,  $C_2H_2$  and  $H_2$  reactant gases which had been pre-heated were flowed through a vertical reactor and then reacted to form the product that condensed at a relative high deposition rate. The product was collected on filter paper at the reactor exit. The reaction formula is



### 2.3. Ample Characterization

The morphology and aggregation situation and size of the microcrystal were observed by TEM (H-600-2, Tokyo, Japan), small amounts of particles were diluted in alcohol for 30 min's process of ultrasonic dispersing, and several drops of the sample were placed on Cu grids. The crystal phase was characterized by method of XRD (D/max-rB, Tokyo, Japan) using CuK $\alpha$  radiation and HR-TEM (JEM-2010, Japan).

## 3. Results and Discussion

### 3.1. XRD spectrogram of WCl<sub>6</sub>

Prepared WCl<sub>6</sub> powders are generally crystals with brown color and with stable properties. Also, it contains a little bit of mauve powders, which changes light yellow for a period of time. Fig.1 shows the XRD spectrograms of the obtained powders. By studying the XRD patterns of the brown part (see Fig.1a), it is confirmed from comparison the interplanar spacing, d, of corresponding peaks with that of JCPDS standard card that this part of product is exactly WCl<sub>6</sub> crystal, no impurities existed. It is known that reaction  $\text{W} + 3\text{Cl}_2 \rightarrow \text{WCl}_6$  takes place at above 600°C . If the Cl<sub>2</sub> and W powder could be provided as hardly contained any impurities, the brown WCl<sub>6</sub> crystal with high purity could be obtained at 600~800°C [11].

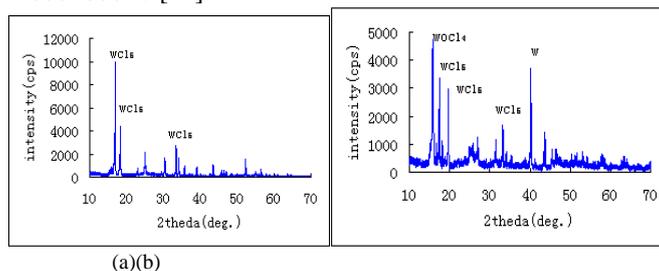
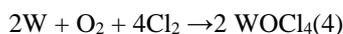
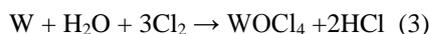
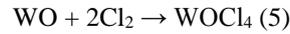


Fig. 1. XRD patterns of the WCl<sub>6</sub> product (a) the brown part and (b) the mixed product

Fig. 1b is the XRD spectra of the very little amount of mauve powders. It can be seen from the spectra that WCl<sub>6</sub>, WOCl<sub>4</sub> and little W powders are contained in this part. Analysis indicates that WOCl<sub>4</sub> is resulted from the impurity of Cl<sub>2</sub> or W powder. If Cl<sub>2</sub> used in experiment contained any water vapor or air, it would bring the reaction



Even slight amount of impurity WO in W powder, it also would result in WOCl<sub>4</sub>



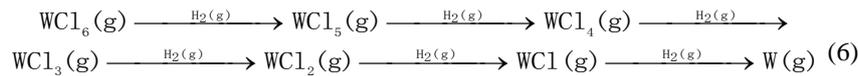
WOCl<sub>4</sub> powders change its color from the initial mauve to light yellow over a period of time. And because of the low boiling point of WOCl<sub>4</sub>, it easily decomposes into WO<sub>3</sub>, which is light yellow, and WO containing slight amount of chlorine.

W powders in the sample come from W that didn't work together in the previous reaction, yet it wouldn't affect the quality of WC. XRD results show that the prepared material is WCl<sub>6</sub> powders containing slight amount of impurity of WOCl<sub>4</sub>.

### 3.2. XRD spectrogram of WC

Fig.2 shows the XRD spectrograms for powders prepared by plasma CVD. Analysis reveals that there are WC, W<sub>2</sub>C and WC<sub>1-x</sub> containing in the product.

Generally, it is thought that WC preparation experienced two steps: the first deoxidize and then carbonize. Detailed dioxide process is illuminated as following



Because of the high temperature of plasma, hydrogen atom derived from H<sub>2</sub> decomposition and further ionization to form H<sup>+</sup>. H and H<sup>+</sup> greatly enhance the deoxidizing ability of H<sub>2</sub>, as well as increase the chemical activity of hydrogen. WCl<sub>6</sub> is vaporized and then is heated by the high temperature of plasma in the reactor, existing as the form of gas molecule instead of solid particles. Gas-gas reduction reaction between highly activated H<sub>2</sub> and gaseous WCl<sub>6</sub> produces highly active W atom. At the high temperature of plasma, W atoms and active carbon atoms from C<sub>2</sub>H<sub>2</sub> collide against each other and combined in reaction. In the process, simple hexagonal WC crystal is formed through optimal grouping of W-C lattice movement, and further WC vapor comes into formation. Saturated WC vapor begins to condense and nucleate through homogenous nucleation when experiencing fast cooling, forming many minute groups, which would grow bigger and bigger if further cooled, finally forming WC nanocrystallites.

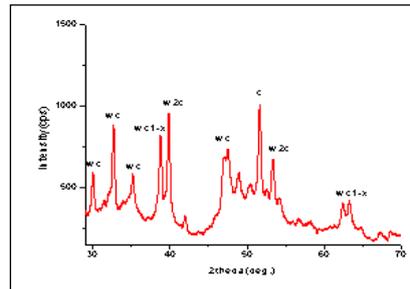


Fig. 2. XRD patterns of the WC product

At temperature as high as 2000K or even higher, different product phases region, such as  $W_2C$ ,  $WC_{1-x}$ , WC, etc, would be distinct by different carbon concentration and reaction temperature. Because carbon concentration and temperature distribution in the reactor are not perfectly even, and synchronously with different reaction time, so different products are obtained under plasma condition. It is hard to prepare WC with high purity, usually, Tungsten carbide of other valence co-existed for the homogeneous phase of WC is really narrow, the influencing factor for WC formation is complicated and the particularly strict requirements for carbon concentration and temperature are needed.

### 3.3. TEM and HR-TEM Analysis of WC

Fig.3 shows the TEM and HR-TEM photographs for WC sample, respectively. It is observed from TEM shown in Fig.3a that most WC particles are ball-shaped and a few of them are stick-shaped or pack-shaped. They appear as aggregated groups with the average diameters of about 50 nm. The corresponding HR-TEM photograph in Fig.3b shows that the particles with the form of aggregated groups are composed of many aggregated nanocrystallites with single, dense structure and 5-15nm in diameter. The reason for aggregation resulted in small particle size with high surface activity, which could aggregate to reduce energy and keep stable states. Fig.3b also indicates that lattice constant,  $a$ , of the microcrystal is about 0.3 nm, which coincides with that of WC lattice constant. All these led the conclusion that the obtained particles are WC nanocrystallites, and the product is the aggregations of WC nanocrystallites. Because the diameter of reactant entrance is smaller than the reactant entry pipe, reactant gases get into the reaction region by eject, getting well mixed. Product gas stream cool down to form particles by water. On the other hand, because the diameter of reaction region is much smaller than that of the cooling region, a sudden broadening exists at the joining point, this resulted the expansion and quencher for the product gas stream, thus preventing WC particles from growing bigger. From the TEM photograph of the few existing stick shape substances as

shown in Fig.3c,it can be seen that they are composed of dense layers, hinting that they are single crystal.

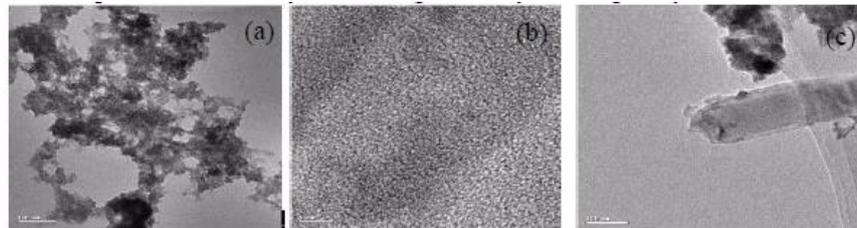


Fig. 3. Particle shape and size distribution of WC nanocrystallites (a) TEM photograph of WC, (b) HR-TEM photograph of WC, (c) TEM photograph for the stick-shaped WC product.

#### 4. Conclusion

By providing  $\text{Cl}_2$  to tungsten powder in silica tube in electric stove at  $600\sim 800^\circ\text{C}$ , purple black and metal-shine  $\text{WCl}_6$  is obtained. Gray black hard WC nanocrystallites are produced with  $\text{WCl}_6$  vapor,  $\text{C}_2\text{H}_2$  and  $\text{H}_2$  as raw materials through RF-PCVD method.

XRD results indicate that the prepared  $\text{WCl}_6$  is of high purity, containing slight amount of mauve impurity  $\text{WOCl}_4$ , which changes into light yellow  $\text{WO}_3$  and  $\text{WO}$  containing minor chloride. It is really difficult to obtain single product because WC powder prepared through PCVD displays a narrow homogeneous phase, and the requirements for temperature and carbon concentration are really rigor. TEM results show that most of WC particles are globular with average diameter about 50nm and a few are stick shape or package shape. HR-TEM results show the lattice constant  $a=0.3\text{nm}$ , which coincides with that of WC crystal, indicating that the particles were all WC nanocrystallites.

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