

Determination of potassium content in potassium hydrogen terephthalate used for fireworks and firecrackers based on Energy Dispersive X-ray Fluorescence Spectrometry (EDXRF)

Wu Jun-yi^{1, a}

¹Technical Center for Dangerous Goods Testing of Guangxi Entry-Exit Inspection and Quarantine Bureau, Beihai, Guangxi, China

^ajohnny-woo@vip.163.com

Keywords: fireworks and firecrackers; potassium content; potassium hydrogen terephthalate; determination; Energy Dispersive X-ray Fluorescence Spectrometry; EDXRF.

Abstract. Methods used for the determination of potassium content in potassium hydrogen terephthalate are mostly based on traditional chemical method, which is lengthy and cumbersome. If inductively coupled plasma emission spectrometry or atomic absorption spectrometry are used to determine the potassium with high content, the sample solution must be highly diluted, and it must produce errors in measurement and calling into question the reliability of the data. The method mentioned in this paper is about the determination of potassium content in potassium hydrogen terephthalate used for fireworks and firecrackers based on energy dispersive X-ray fluorescence spectrometry by controlling matrix effects between elements. Using sample solution of potassium hydrogen terephthalate in specific concentrates, the potassium content can be determined by the specific calibration curve established with an intensity calibration. This method can provide high accuracy and good precision in a short time with a simple process by efficiently controlling the matrix effects. It can fully meet the requirements for the determination of potassium in potassium hydrogen terephthalate used for different kinds of fireworks and firecrackers around the world, and it has good generalization and practicability. The average recovery of the method can be 96.49%~99.63%, allowing for a difference of 0.5%.

Introduction

In China, fireworks and firecrackers are very important consumer recreational products in people's everyday life since ancient times. Gorgeous colors produced by fireworks and firecrackers are even the potassium ing role of foiling festal atmosphere in every grand holiday celebrations. In recent years, with the rapid development of global trade, fireworks and firecrackers are becoming more and more popular all over the world, more and more consumers are fascinated by different kinds of patterns, pictures, and sound effects of fireworks and firecrackers. Potassium hydrogen terephthalate is the most popular chemical material used to be oxidant in fireworks and firecrackers. Quantitative analysis method of the potassium content reported in current literature is limited to traditional chemical analysis, such methods have the following disadvantages: (1) Long detecting period. Generally, it will take a skilled technician two whole days or so to complete the detection. (2) The operation is more complicated. It needs to go through many steps such as dissolving sample, filtration, precipitation collection, drying and weighing precipitation and ect. Comparing with traditional chemical analysis methods, this method based on energy dispersive X-ray fluorescence spectrometry (EDXRF) has the advantages of simple operation steps, short period of detection, high accuracy and good precision.

Theory

Concept of the method: considering the weight of the sample is 1.0g, constant volume is 0.5L and the concentrations of the potassium would be controlled in 0.30g/L~0.40g/L in sample solutions. And it can prove that when the potassium content in the solution is in the range of 0.16 g/L~0.99g/L and the chlorine elements with the content of 0.18 g/L~1.10g/L, there would be little matrix effects among elements. So we can establish a working curve which contains the potassium elements with the content of 0.16 g/L~0.99g/L and the chlorine elements with the content of 0.18 g/L~1.10g/L to

determine the potassium content in the sample solution. In accordance with the relevant safety regulations, the sample was ground into powder of less than 180 micron. Then the sample powder is placed in an oven at 50°C–55°C and dried for 4 hours, and then placed into a dryer for cooling down to room temperature. Pretreated sample is fully dissolved in 150mL pure water and then filtered into volumetric flask as sample solution. The sample solution can be put into the sample cup and placed in the tank of the EDXRF to measure the fluorescence intensity of the potassium elements. The actual content of potassium element in the sample can be calculated from the concentrations of the potassium reading by the working curve.

Experiment section

Reagents

Unless otherwise stated, all the reagents should be guaranteed reagents and pure water is secondary grade water as described in ISO 3696(1987). Nitric acid (1+4): mix nitric acid and pure water thoroughly according to the proportion of 1:4. Nitric acid (1+19): mix nitric acid and pure water thoroughly according to the proportion of 1:19. Standard working solution of the potassium chloride : Weigh 3.5 g potassium chloride reference materials (accuracy to 0.1 mg), and put it in a 300 ml beaker, add 150 mL pure water, heat the beaker and make the sample solution slightly boiling on an electric stove for 10 min. After the solution is cool down to the room temperature, transfer the solution into a 500 ml volumetric flask and add 10 ml nitric acid (1+4) and pure water to the scale. Then we can separately pipette the standard working solution of the potassium chloride with volume 5mL、10mL、15mL、20mL、25mL and 30mL into six 100 mL volumetric flasks, and add nitric acid (1+19) to reach 100 mL in each volumetric flask, mix thoroughly for later use. Concentrations of the standard working solution in different flasks are shown in Table 1.

Table 1 Concentrations of series standard working solutions of potassium
g/L

element	concentrations of series standard working solutions					
	N1	N2	N3	N4	N5	N6
K	0.1658	0.3316	0.4974	0.6632	0.8289	0.9947
Cl	0.1828	0.3657	0.5485	0.7314	0.9142	1.0970

Instrument and apparatus

Oven with accuracy to $\pm 2^{\circ}\text{C}$. Analytical balance with accuracy to 0.1 mg. Energy dispersive X-ray fluorescence spectrometer (EDXRF): The United States Thermo Fisher (former Thermo Electron Corporation) Company QUANT'X series.

Operation step

- (1) Weigh the sample of about 1.0 g, accuracy to 0.1 mg.
- (2) Place the sample into a clean 300 mL beaker, add 150 mL pure water into the beaker, then place the beaker on an electric stove to make the solution boiled for about 10 min. Filter the solution through filter paper to an 0.5L volumetric flask, wash the beaker and the filter paper several times with pure water, and add 20mL Nitric acid (1+4) and make a constant volume with pure water after the filtered solution cooling down to room temperature.
- (3) Parameters of the EDXRF instrument parameters are shown in Table 2.

Table 2 Parameters of the EDXRF instrument

Filter	Fiber
Collimator	8.8mm
Voltage	12v
Electric current	Auto
Analysis time	35s
Count rate	Medium
Atmosphere	Air
Matrix effects	Not considered
Energy range	0~20kev
Analysis technique	Intensity correction
sample thickness	$\geq 15\text{mm}$

(4) Calibration (working) curve: according to the requirements of the method and the instrument criteria, we set the instrument to optimum analysis conditions, and adjust it to the best working condition, and determine spectral intensity of the series standard solution from N_1 to N_6 to establish the calibration (working) curve with the elemental concentrations as independent variable and the spectral intensity as the dependent variable. The linear correlation coefficient of the regression curve should be 0.99 or higher.

(5) Sample determination: determine the fluorescence intensity of the potassium in blank solution and every sample solution under the best analysis condition and read the concentrations from the calibration curve according to the spectral intensity.

Results calculation

Content of the potassium element in the sample can be calculated as mass fraction W and its value shown in% according to the following formula.

$$\omega = \omega_0 \times \frac{1.0}{m} \times \frac{V}{500}$$

Where: ω_0 —the content of the potassium in the sample read by the working curve, expressed in %.
 m —quantity of the sample, expressed in milligrams (g).

V —constant volume of the volumetric flask used for the sample solution, expressed in liters(mL).

ω — the content of the potassium in the sample, expressed in %.

1.0—assume that quantity of the sample, expressed in milligrams (g).

500—assume that constant volume of the volumetric flask used for the sample solution, expressed in liters(mL).

Results and discussion

Selection of standard solution.

Considering that all the potassium element come from potassium hydrogen terephthalate. In order to make the standard solution as consistent as possible with the sample solution, the K-Cl standard solution would be selected to make the working curve. It proved that when the concentration of the potassium element is controlled to the range of 0.16 g/L~0.99g/L and the chlorine element with the content of 0.18 g/L~1.10g/L, it would have little matrix effects on the potassium element and can be basically ignored. Because the contents of other impurity elements such as sodium and manganese

are all mostly less than the chlorine element, they would also have little matrix effects on the potassium element in the sample solution.

Recovery test

To assess the accuracy of the method, we used different standard substances and serious potassium standard solutions to do the recovery test. We dissolved the samples and determined the contents of potassium in the sample solution. The values of the potassium contents we measured are compared with the theoretical ones, and the experimental data is shown in Table 3 below.

Table 3 Recovery test results

NO.	Reference materials	Reference Code	Potassium content of nominal (%)	Potassium content of measurment (%)	recovery rate (%)
1	Potassium hydrogen terephthalate	K01	17.18	16.87	98.20
2	Potassium hydrogen terephthalate	K02	16.91	16.65	98.46
3	Potassium hydrogen terephthalate	K03	16.73	16.49	98.57
4	Potassium standard solution	K04	7.12	6.87	96.49
5	Potassium standard solution	K05	11.41	11.11	97.37
6	Potassium standard solution	K06	14.25	13.94	97.82
7	Potassium standard solution	K07	21.38	21.12	98.78
8	Potassium standard solution	K08	38.5	38.29	99.45
9	Potassium standard solution	K09	57.76	57.41	99.39
10	Potassium standard solution	K10	61.44	61.21	99.63
Average (X)					98.42
standard deviation (S)					0.94

Conclusions

Energy dispersive X-ray fluorescence spectrometry (EDXRF) is used to determine the potassium content in potassium hydrogen terephthalate used for fireworks and firecrackers, this method is accurate and quick with high accuracy and good precision. The recovery is 96.49%~99.63%.The allowable differential value was 0.5% between two single tests under repeatable conditions. In other word, this method can completely satisfy the requirements of the fireworks and firecrackers industry.

References

- [1] Leif, H. C.; Allan, A. Determination of sulfur and heavy metals in crude oil and petroleum products by energy-dispersive x-ray fluorescence spectrometry and fundamental parameter approach. *Anal. Chem.*,53(12),pp.1788-1792,1981.
- [2] Fei, H; Pierre, J. V. E. General approach for quantitative energy dispersive x-ray fluorescence analysis based on fundamental parameters. *Anal. Chem.*,63(20),pp.2237-2244,1991.
- [3] Standard test method for determination of lead in paint layers and similar coatings or in substrates and homogenous materials by energy dispersive x-ray fluorescence spectrometry using multiple monochromatic excitation beams,ASTMF2853,American Society for Testing and Materials Publications, 2010.
- [4] Standard test method for analysis of uranium and thorium in soils by energy dispersive x-ray fluorescence spectroscopy, ASTM C1255,American Society for Testing and Materials Publications, 2011.
- [5] Duan Tiyu, Li Suqing, et al. Jewellery,Determination of precious metal content - Method using x-ray fluorescence spectrometry,China National Recommended Standard: GB/T 18043-2013, China Standards Press: Beijing, 2014.
- [6] Leoni. L,Saitta. M, X-ray fluorescence analysis of powder pellets utilizing a small quantity of material, *X-ray Spectrom*, 3,pp.74-77,1974.
- [7] Rose W.I.,Bornhorst T.J.,Sivonen S.J., Rapid high-quality major and trace element analysis of powdered rock by x-ray fluorescence spectrometry, *X-ray Spectrom*, 15,pp.55-60,1986.
- [8] Gy,Pierre M,The analytical and economic importance of correctness in sampling,*Anal.Chim.Acta*, 190,pp.13-23,1986.
- [9] ZHOU Tong-hui ,WANG Er-kang,LU Wan-zhen ect. Handbook in analytical chemistry(second edition), basic knowledge of and safety knowledge(the first volume) , China Chemical industry press:pp.568-580,1997.

邮寄地址：广西北海市海城区贵州路41号 吴俊逸 13877991366

邮编：536000